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Chapter 4 Atomic  
Structure 33 SECTION  
4.1 DEFINING THE  
ATOM (pages 101–103)

This section describes  
early atomic theories  
of matter and provides  
ways to understand the  
tiny size of individual  
atoms. Early Models of  
the Atom (pages  
101–102)

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Name Date Class

**ATOMIC STRUCTURE**

**4**

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**Ch 4: Atomic  
Structure**

Section 4.2 Structure of  
the Nuclear Atom One  
change to Dalton's  
atomic theory is that  
atoms are divisible into  
subatomic particles:

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Electrons, protons, and neutrons are examples of these fundamental particles. There are many other types of particles, but we will study these three.

**Chapter 4 Atomic  
Structure - Henry  
County School  
District**

Chapter 4 Atomic  
Structure Section 4.3  
Modern Atomic Theory  
(pages 113-118) This  
section focuses on the

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arrangement and  
behavior of electrons in  
atoms. Reading  
Strategy (page 113)  
Sequencing After you  
read, complete the  
description in the flow  
chart below of how the  
gain or loss of energy  
affects electrons in  
atoms. For

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4.3 Modern Atomic  
Theory**

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*Page 8/27*



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4.05 Atomic Number,  
Mass Number, Isotopes  
and Isobars 4.05

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number, Isotopes and  
Isobars. Chapter 4  
Structure of Atoms -  
Test. 05 The  
Fundamental Unit of  
Life 5 Topics | 6  
Quizzes 5.01  
Introduction to cell  
5.01 Introduction to  
Cell. 10.03 Plasma  
Membrane 5.02 Plasma  
Membrane ...

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Class 9 Chapter 4  
Structure of Atom ...**

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The Structure of an  
Atom 4.3: Modern  
Atomic Theory

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He recognized a set of patterns that applied to all elements. 2. The atomic mass of an element is the average mass of all the isotopes of that element. 3. Patterns appeared when he arranged the elements in order of increasing atomic mass. 4. true 5. three 6. in a regular,

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Repeated pattern 7.  
periodic table 8. atomic  
number 9.  
Answers

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CHAPTER 4 TEST:  
Atoms, Atomic Theory  
and Atomic Structure a  
student of Thompson,

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who tested the current theory of the atomic structure, using alpha particles (positively charged) which are directed at a sheet of gold foil, according to thompson's theory the particles should have passed through the gold with only slight deflection due to the positive charge in the foil, but he proved this wrong because majority of the particles passed

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through the gold...

**[Book] Chapter 4  
Atomic**

Chapter 4 Atomic  
Structure Section 4.2  
The Structure of an  
Atom (pages 108–112)  
This section compares  
the properties of three  
subatomic particles. It  
also discusses atomic  
numbers, mass  
numbers, and isotopes.  
Reading Strategy  
(page 108) Monitoring  
Your Understanding



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**Chapter 4 Atomic  
Structure Section  
4.2 The Structure of  
an ...**

This unit's focus is on atomic structure. Students will cover concepts related to atomic mass, isotopes, electron configurations and periodicity.

Chapter 4 - Atomic Structure

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**Unit 2 - Chapters 4,  
5 & 6 - Mrs. Gingras'  
Chemistry Page**

This product is a  
PowerPoint that covers  
the basics of the  
structure of an atom. It  
covers atoms as the  
building blocks of  
matter, protons,  
neutrons, electrons  
and their location  
within the atom, a brief  
explanation of electric  
charges, and the  
charges of the  
subatomic particles,

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and finally the concept  
of how the numbers of  
subatomic particles  
makes an atom unique  
from any other type of  
atom ...

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Structure, atom.

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atomic mass. atomic  
mass unit (amu)  
atomic number (Z) the  
smallest unit of an  
element that maintains  
the properties.... the  
weighted average of  
the masses of the  
isotopes of an eleme....  
a unit of mass equal to  
one-twelfth the mass of  
a carbon-12 at....  
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IPOD # 7 - atomic  
structure chart Review  
HW Worksheet: Atomic  
Structure, Sheet 1  
(finish) Chapter 4  
Notes, Slides 14-16:  
Effect of changing  
particles Chapter 4  
Notes, Slide 17: Isotope  
definition Chapter 4  
Notes, Slide 18-20:  
Practice Lab: Beans  
HW: Worksheet:  
Chapter 4 - Atomic

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Model Timeline. Day 3 -  
IPOD # 8 - isotopes FC  
- Atomic ...  
Answers

**McLaughlin,  
Kimberly / Atomic  
Structure**

Name: \_\_\_\_\_ Date: \_\_\_\_\_

Period: \_\_\_\_\_ Chapter 4:

History of the Atom

Guided Notes

Democritus Years

Alive: \_\_\_\_\_ • There are  
various basic \_\_\_\_\_  
from which all matter is  
made. • Everything is  
composed of small

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\_\_\_\_\_ moving in a \_\_\_\_\_.

- Some atoms are round, pointy, oily, have hooks, etc. to account for their \_\_\_\_\_

- Ideas rejected by leading philosophers because void = \_\_\_\_\_  
Democritus ...

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