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[2] (b) Marit arranged the values she found in Table 12 into an energy cycle.
$$\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$$
$$3 \text{O}_2 \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}$$
$$-141 \text{ kJ mol}^{-1} \quad -286 \text{ kJ mol}^{-1} \quad -1560 \text{ kJ mol}^{-1}$$

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(hydrogenation) (l) Calculate the value for the enthalpy of hydrogenation of ethene from the energy cycle.....

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8 20 2 2 2 2 o correct reactants and products; [2] correct coefficients; M2 can only be scored if M1 correct. (b) Local pollutant: carbon monoxide/CO / volatile organics/VOCs / nitrogen oxide/NO / (unburnt) hydrocarbons; Do not accept methane/CH₄, ethane/C₂H₆, propane/C₃H₈ or butane/C₄H₁₀. Global pollutant: nitrogen oxide/NO / carbon ...

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$\text{CH}_3\text{CHOH}(\text{l}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O}(\text{l})$ $\Delta H = -1367 \text{ kJ mol}^{-1}$
 $\text{C}_8\text{H}_{18}(\text{l}) + 12\frac{1}{2}\text{O}_2(\text{g}) \rightarrow 8\text{CO}_2(\text{g}) + 9\text{H}_2\text{O}(\text{l})$ $\Delta H = -5470 \text{ kJ mol}^{-1}$
(i) Calculate the mass, in g, of ethanol and octane in 1.00 kg of the fuel mixture. [1].....

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$(\text{CH}_3)_2\text{CHCH}_2\text{OH}$; $(\text{CH}_3)_3\text{COH}$; [2 max]

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Accept condensed or full structural formulas. Penalise missing H atoms or incorrect bonds (such as C-HO, C-H 2C) only once in the whole paper. (c) C H CHOCH [O] C H COCH H O25 3 25 3 2+→ +; [1] Accept condensed or full structural formulas. Accept [O] on top of the arrow. Do not accept equation without H 2O.

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